

Daily Tutorial Sheet 5

Level – 1

- 61.(A)** Fluoro apatite is $3\text{Ca}_3(\text{PO}_4)_2 \cdot \text{CaF}_2$.
- 62.(D)** Order of bond dissociation enthalpy $\text{Cl}_2 > \text{F}_2 > \text{Br}_2 > \text{I}_2$
- 63.(B)** conc. HNO_3 is yellow coloured liquid due to presence of dissolved NO_2 .
- 64.(D)**
- 65.(D)** Order of reactivity of halogen is $\text{F}_2 > \text{Cl}_2 > \text{Br}_2 > \text{I}_2$.
- 66.(B)** $\text{PCl}_5 + \text{H}_2\text{O} \longrightarrow \text{POCl}_3 + 2\text{HCl}$
(A) (B)
- $\text{POCl}_3 + 3\text{H}_2\text{O} \longrightarrow \text{H}_3\text{PO}_4 + 3\text{HCl}$
(B) (C)
- 67.(D)** $\ddot{\text{P}}\text{H}_3 + \text{H}^+ \longrightarrow \text{PH}_4^+$
Bond angle of PH_3 is almost 90° . So, lone pair of P resides in almost pure s-orbital.
- 68.(A)** F^- can-not be oxidized by MnO_2 because F_2 is strongest oxidizing agent.
- 69.(D)** He is used in gas-cooled nuclear reactors.
- 70.(A)** Xenon is most easily liquefiable gas because of stronger intermolecular forces.
- 71.(D)** Metals that not reacts with most of the reagent are called noble metals.
- 72.(C)** All orbitals of inner shell and valence shell are filled and having $ns^2 np^6$ configuration for valence shell.
- 73.(A)** All the noble gases (except He) are adsorbed on the surface of coconut charcoal. Being light, He is not adsorbed. Ne is also less adsorbed on coconut charcoal.
- 74.(C)** The electrical resistance of He is zero and it is used as super conducting liquid.
- 75.(D)** Noble gases are adsorbed by activated coconut charcoal.